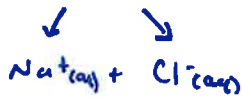


4/9

Electrolyte


1 mol NaCl(aq)



Amt. of NaCl: 1 mol

Amt of solute particles: 2 mol

$$i = \frac{2}{1} = 2$$

Halogen  
  
 "aryl halides"  
 oily dispersion dominated  
 denser than H<sub>2</sub>O

Non Electrolyte

1 mol sucrose C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>

Amt of sucrose: 1 mol

Amt of solute particles: 1 mol

$$i = \frac{1}{1} = 1$$

200g NaCl(aq)

Amt. of NaCl:  $200\text{g} \left( \frac{\text{mol}}{58.44\text{g}} \right) = 3.422 \text{ mol}$

Amt solute particles:  $200\text{g} \left( \frac{\text{mol NaCl}}{58.44\text{g}} \right) \left( \frac{2 \text{ mol part.}}{\text{mol NaCl}} \right) = 6.844 \text{ mol solute particles}$

Strong Electrolyte

Strong Acids

Strong Bases

Soluble salts

Weak Electrolytes

Weak acids

Weak bases

alcohols

Distilled water H<sub>2</sub>O

Tap water H<sub>2</sub>O

Sodium chloride NaCl

Sugar

Vinegar CH<sub>3</sub>COOH

Hydrochloric Acid HCl

Sodium Hydroxide NaOH

Ethanol C<sub>2</sub>H<sub>5</sub>OH

Barium Sulfate BaSO<sub>4</sub>

Electrolytes Canvas Exam