

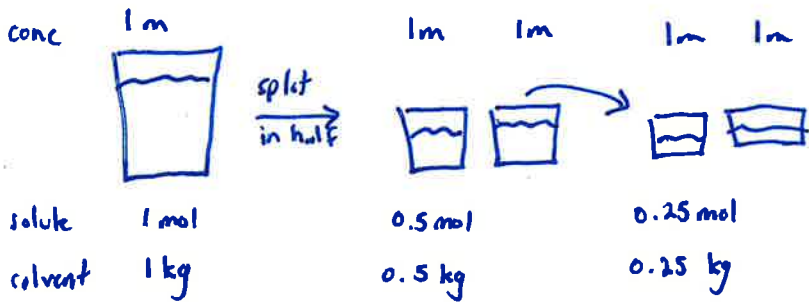


mm. 58.44 g mol⁻¹



$V_{\text{soln}} = ?$

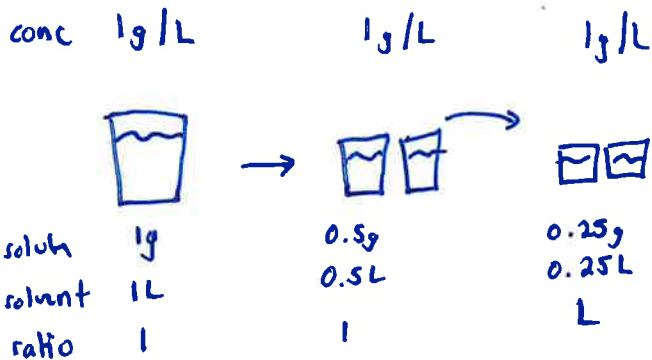
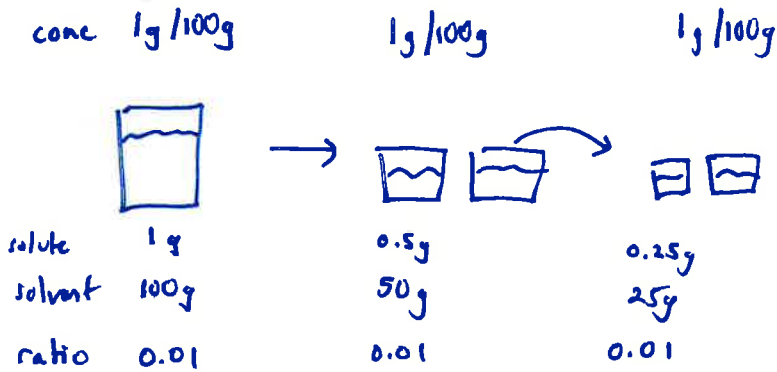
- ⊕ mass of soln
- ⊕ density of soln
- ⊖ 1 and 2 to get V of



"RATIOS"

How to change concentration?

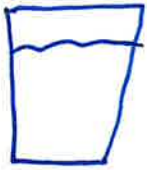
- ① More concentrated soln
 - Remove solvent
 - Add solute
- ② More dilute soln
 - Remove solute
 - Add solvent



Fin Exam
20 Qs 2.25/Q
45 min
12 M.C.

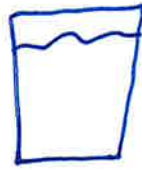
Q6 HW Exam: 10
solns → Read
Like Decolour litmus paper → saturation of solubility units

conc. 1% by mass



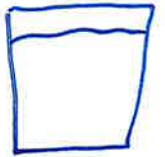
solute	1g	$\frac{1g}{100g} \times 100\% = 1\%$
solvent	99g	
solution	100g	

1%



solute	0.1g	$\frac{0.1g}{10g} \times 100\% = 1\%$
solvent	9.9g	
solution	10g	

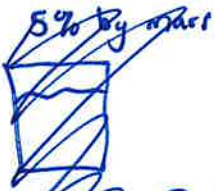
1%



xg

$$\frac{xg}{10g} \times 100\% = 1\%$$

so



What is mass of solution?