

molality

m

$$\frac{\text{mol solute}}{\text{kg solvent}}$$

molarity

M

$$\frac{\text{mol solute}}{\text{L solution}}$$

mass concentration

P

$$\frac{\text{g solute}}{\text{L solution}}$$

g/100g

g/100g

$$\frac{\text{g solute}}{100 \text{ g solvent}}$$

mole fraction

X

$$\frac{n_{\text{solute}}}{n_{\text{solution}}}$$

mole percent

mol %

$$\chi_{\text{solute}} \times 100\%$$

mass fraction

3

$$\frac{m_{\text{solute}}}{m_{\text{solution}}}$$

mass percent

mass %

wt%, %wt, w/w%, wt/wt%

$$\omega_{\text{solute}} \times 100\%$$

mass volume percent

m/v%

w/v%

$$\frac{m_{\text{solute}} \text{ (g)}}{V_{\text{solution}} \text{ (mL)}} \times 100\%$$

parts by mass

pp_x

$$\frac{m_{\text{solute}}}{m_{\text{solution}}} \times \text{multiplication factor}$$

parts by volume

pp_x

$$\frac{V_{\text{solute}}}{V_{\text{solution}}} \times \text{multiplication factor}$$